Neuroinformatics 2016

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Membrane potential and conductance-based model

Cell membrane

- Phospholipidic bilayer
- Ion channels



Membrane potential: Diffusion

- Permeable membrane + 2 solutions with different ion concentration
- lons diffuse until equilibrium is achieved (concentration + electrical)



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Nernst potential



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Calculation of Nernst potential

Nernst calculation for one ion

$$V = \frac{RT}{zF} \log \frac{[c_{out}]}{[c_{in}]}$$

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V - membrane potential [V = J/C]

[cin] - concentration inside the cell $[mol/m^3]$

[cout] - concentration outside the cell $[mol/m^3]$

- R gas constant [J/K/mol]
- T temperature [K]
- F Faraday's constant [C/mol]

Membrane simulation

$$I_{C}(t) = C_{m} \frac{dV}{dt}(t)$$

$$\tau \frac{dV}{dt} = V_{Cl} - V(t) + \frac{I_{stim}(t)}{Ag_{Cl}}$$

$$\tau = \frac{C_{m}}{g_{Cl}}$$

$$I_{stim} \rightarrow 0$$

$$I_{stim} \rightarrow 0$$

$$I_{stim} \rightarrow 0$$

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Membrane simulation



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Real membrane: multiple ions



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lon	Intracellular c.	Extracellular c.	Equilibrium Potential
Na+	15 mM	145 mM	$V_{Na} = +60 mV$
K+	150 mM	4 mM	$V_{K} = -97 mV$
CI	10 mM	110 mM	$V_{Cl} = -64mV$
Ca2+	70 nM	2 mM	$V_{Ca} = +137 mV$
H+	63 nM (pH 7.2)	40 nM (pH 7.4)	$V_{H+} = -12mV$

Multiple ions: Goldman equation

$$V = \frac{RT}{zF} \log \left(\frac{P_{Na}[Na^{+}]_{out} + P_{K}[K^{+}]_{out} + P_{CI}[CI^{-}]_{in}}{P_{Na}[Na^{+}]_{in} + P_{K}[K^{+}]_{in} + P_{CI}[CI^{-}]_{out}} \right)$$

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V - membrane potential [V = J/C] $[M]_{in}$ - concentration of ion M inside the cell $[mol/m^3]$ $[M]_{out}$ - concentration of ion M outside the cell $[mol/m^3]$ $[P]_M$ - Membrane permeability for ion M [m/s]

Further Readings

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